

## Carbonate hardness C 20

Test kit for the determination of p and m value and carbonate hardness

## Method:

Acid / base titration

## Contents of test kit (\*refill pack):

sufficient for approx. 200 tests with an average alkalinity of 10 °d or 3.6 mmol/L HCl  
 10 mL indicator p\*  
 10 mL indicator m\*  
 100 mL titration solution TL C 20\*  
 1 test tube with 5 mL ring mark  
 1 titration syringe 0–20 °d resp. 0–7.2 mmol/L HCl  
 (1 graduation mark  $\Delta$  0.5 °d resp. 0.2 mmol/L)  
 2 plastic dropping tips

## Hazard warning:

Information regarding safety can be found on the box' label and in the safety data sheet. You can download the SDS from [www.mn-net.com/SDS](http://www.mn-net.com/SDS).

## Procedure:

- Rinse test tube several times with the test sample and fill to ring mark.
- Determination of p value (p alkalinity):  
Add 1 drop of indicator p and mix by shaking. If test sample remains colorless, the p value is 0. If test sample turns red, proceed as follows:
- Put dropping tip onto the titration syringe, press down the plunger, dip the tip into the titration solution TL C 20 and draw up plunger slowly, until the lower rim of the black plunger O-ring agrees with value 0 on the barrel scale. The small air pocket below the plunger tip does not disturb the determination.
- Addition of the titration solution: We recommend taking the syringe in the left hand and the test tube in the right hand (see drawing) and adding titration solution dropwise while smoothly shaking the test tube. As soon as the red color vanished completely, read off p value in °d or mmol/L HCl from the syringe barrel (lower rim of the black plunger O-ring) and write down.
- Determination of m value (m alkalinity):  
Add to the remaining water sample in the test tube 1 drop of indicator m and mix by shaking. If the water turns red, the m value is identical to the p value. If the water turns blue, determine the m value as follows:
- Continue titrating with the same syringe filling as used for determination of the p value, until the color turns red. Read off the m value in °d or mmol/L HCl from the syringe barrel (lower rim of the black plunger O-ring) and write down. After addition of a further drop titration solution the color should not change. If the 1st syringe filling is not sufficient to reach color change, fill up the syringe once more with solution TL C 20 and titrate to color change as described before. Add the additional used syringe content to the value written down.
- Carbonate hardness:  
Normally the m value is identical with the carbonate hardness (mmol/L HCl or °d). Attention, do not mix up mmol/L HCl with mmol/L CaO\*.  
If this results in a carbonate hardness, which is higher than the total hardness, the result is discarded. In this case the carbonate hardness is equivalent to the total hardness.

°d	°e	°f	mg/L CaO	mg/L CaCO <sub>3</sub>	mmol/L H <sup>+</sup>
1	1.3	1.8	10	18	0.36
2	2.5	3.6	20	36	0.71
3	3.8	5.4	30	54	1.07
4	5.0	7.1	40	71	1.43
5	6.3	8.9	50	89	1.78
6	7.5	10.7	60	107	2.14
7	8.8	12.5	70	125	2.50
8	10.0	14.3	80	143	2.86
9	11.3	16.1	90	161	3.21
10	12.5	17.8	100	178	3.57

This method can be used also for analyzing sea water.

## Disposing of the samples:

Information regarding disposal can be found in the safety data sheet. You can download the SDS from [www.mn-net.com/SDS](http://www.mn-net.com/SDS).

## Remarks:

- m > 2p → carbonate hardness = m; sample contains CO<sub>3</sub><sup>2-</sup>, HCO<sub>3</sub><sup>-</sup>  
 m = 2p → carbonate hardness = m; sample contains CO<sub>3</sub><sup>2-</sup>  
 p < m < 2p → carbonate hardness = 2 (m - p); sample contains CO<sub>3</sub><sup>2-</sup>, OH<sup>-</sup>  
 m = p → carbonate hardness = 0; sample contains OH<sup>-</sup>  
 p = 0 → carbonate hardness = m; sample contains CO<sub>2</sub>, HCO<sub>3</sub><sup>-</sup>

